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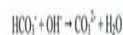


so many fake sites. this is the first one which worked! Many thanks

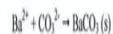
This procedure involves two titrations. First, total alkalinity (moles of bicarbonate + moles of carbonate) is measured by titrating the mixture with standard HCl to a methyl orange end point:



A separate aliquot of unknown is treated with excess standard NaOH to convert HCO_3^- to CO_3^{2-} :



Then all the carbonate is precipitated with BaCl_2 :



The excess NaOH is immediately titrated with standard HCl to determine how much HCO_3^- was present. From the total alkalinity and moles of bicarbonate, you can calculate the original moles of carbonate present.

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