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NAME _____ DATE _____ CLASS _____

CHAPTER 13 STUDY GUIDE

Section 13.4 Neutralization
 In your textbook, read about neutralization and titration.

For each item in Column A, write the letter of the matching item in Column B.

| | |
|---|--|
| <p>Column A</p> <ol style="list-style-type: none"> A chemical dye that changes color based on the pH of a solution. A method for using a neutralization reaction to determine the concentration of a solution. A reaction in which an acid and a base react to produce a salt and water. A solution of known concentration. An ionic product of an acid-base reaction. The point in a titration in which an indicator changes color. The stoichiometric point of a titration. | <p>Column B</p> <ol style="list-style-type: none"> acid-base indicator end point equivalence point neutralization salt standard solution titration |
|---|--|

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

| Acid | Base | Salt formula | Salt name |
|---------------|------------|--------------|--------------------|
| 8. HCl | KOH | KCl | potassium chloride |
| 9. H_2SO_4 | $Mg(OH)_2$ | | |
| 10. H_3PO_4 | NaOH | | |
| 11. HNO_3 | $Fe(OH)_2$ | | |
| 12. H_2PO_4 | $Ca(OH)_2$ | | |

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

Sequence of Steps

- Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- Fill a buret with the standard solution.
- Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

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