

Download File PDF Electrochemistry Multiple Choice Questions Answers And Explanations

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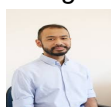
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Electrochemistry - Multiple Choice Questions

1. Which of these chemical reactions is an oxidation-reduction reaction?
A. $\text{Fe} + \text{S} \rightarrow \text{FeS}$ B. $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$ C. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$ D. $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
2. What happens to the oxidizing agent in an oxidation-reduction reaction?
A. It is oxidized and gains electrons. B. It is reduced and it gains electrons. C. It is oxidized and it loses electrons. D. It is reduced and it loses electrons.
3. In which substance does boron have an oxidation number of +3?
A. B_2O_3 B. B_2H_6 C. BH_3 D. HBO_2
4. Which statement is true for an electrochemical cell?
A. Oxidation occurs at the anode only.
B. Reduction occurs at the anode only.
C. Oxidation occurs at both the anode and cathode.
D. Reduction occurs at both the anode and cathode.
5. Given the equation: $2\text{O}^{2-}(\text{aq}) + 2\text{H}^+(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$, which is the correct reduction half-reaction?
A. $\text{O}^{2-}(\text{aq}) \rightarrow \text{O}(\text{g}) + 2\text{e}^-$ B. $\text{O}^{2-}(\text{aq}) + 2\text{e}^- \rightarrow \text{O}(\text{g})$ C. $\text{H}^+(\text{aq}) \rightarrow \text{H}(\text{g}) + 2\text{e}^-$ D. $\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}(\text{g})$
6. What is the E for an electrochemical cell with the following reaction? $2\text{Ag}^+(\text{aq}) + \text{Sn}^{2+}(\text{aq}) \rightarrow 2\text{Ag}(\text{s}) + \text{Sn}^{4+}(\text{aq})$
A. -1.22 V B. 1.26 V C. 1.22 V D. 3.84 V
7. An electrolytic cell is plated with zinc metal in the opposite basin. Which of the following is true?
A. It is a voltaic cell and the reaction is spontaneous.
B. It is a voltaic cell and the reaction is non-spontaneous.
C. It is an electrolytic cell and the reaction is spontaneous.
D. It is an electrolytic cell and the reaction is non-spontaneous.
8. Which of the oxidation states of iodine in the ions IO_3^- and IO_4^- respectively?
A. +4 and -6 B. +4 and -5 C. +5 and -6 D. +5 and -4
9. Which one of the following reactions is a redox reaction?
A. $\text{Fe}^{2+}(\text{aq}) + \text{Cl}_2(\text{g}) \rightarrow \text{FeCl}_3(\text{s})$ B. $\text{NaOH}(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
C. $\text{H}_2\text{SO}_4(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{Na}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l})$
D. $2\text{HNO}_3(\text{aq}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow 2\text{H}_2\text{SO}_4(\text{aq}) + \text{HNO}_3(\text{aq})$
10. Complete the following unbalanced redox equation:
 $\text{Cr}_2\text{O}_7^{2-}(\text{aq}) + \text{Fe}^{2+}(\text{aq}) \rightarrow \text{Cr}^{3+}(\text{aq}) + \text{Fe}^{3+}(\text{aq})$
Which of the following sets of numbers will balance the reaction?
A. 1, 14, 1, 2, 7 B. 3, 1, 3, 2, 7 C. 3, 1, 6, 3, 2, 6 D. 3, 1, 3, 3, 2, 6
11. In which of the following does sulfur have an oxidation number of +7?
A. H_2SO_4 B. SO_2 C. H_2SO_3 D. H_2S
12. What happens to the reducing agent in an oxidation-reduction reaction?
A. It is oxidized and it gains electrons. B. It is reduced and it gains electrons.
C. It is oxidized and it loses electrons. D. It is reduced and it loses electrons.
13. What is the term for the electrode where oxidation occurs?
A. anode B. cathode C. oxidizing agent D. reducing agent

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