

# Download File PDF Pogil Answers Naming Molecular Compounds

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16. For each of the following compounds, indicate whether or not your naming rules from Question 15 will apply. If not, explain why the naming rules do not apply.

17. Using the rules you developed in Question 15, name each of the following molecular compounds.

Molecular Formula	Molecule Name
PBr <sub>3</sub>	phosphorus tribromide
SCl <sub>2</sub>	sulfur dichloride
N <sub>2</sub> F <sub>4</sub>	dinitrogen tetrafluoride
SO <sub>2</sub>	sulfur dioxide
BF <sub>3</sub>	boron trifluoride

18. Write molecular formulas for the following compounds.

Molecular Formula	Molecule Name
S <sub>2</sub> F <sub>10</sub>	Disulfur decafluoride
CCl <sub>4</sub>	Carbon tetrachloride
O <sub>3</sub>	Oxygen trioxide
P <sub>4</sub> O <sub>10</sub>	Tetraphosphorus decoxide
P <sub>4</sub> S <sub>3</sub>	Tetraphosphorus trisulfide

*Handwritten notes:*  
16. No, this is an ionic compound, the positive and negative charges are not in the same group.  
17. Yes, both elements are nonmetals. No, all elements are nonmetals, but there are three elements.  
18. No, all elements are nonmetals, but there are three elements.

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