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Name _____ period _____ AP Chemistry Unit 7 worksheet

1. What is energy? *The measure of the ability to do work*
2. What is heat? *A form of energy that is transferred because of temperature differences*
3. What is the SI unit for heat? *Joule*
4. What is an exothermic reaction? What is the sign of ΔH for an exothermic reaction?
Heat is released, negative
5. What is an endothermic reaction? What is the sign of ΔH for an endothermic reaction?
Heat is absorbed; positive
6. Consider the following reaction, which occurs at room temperature and pressure:
 $2\text{C}(s) \rightarrow \text{C}_2(g) \Delta H = +241.4 \text{ kJ}$
Which has higher enthalpy under these conditions, $2\text{C}(s)$ or $\text{C}_2(g)$?
7. Consider the following reaction:
 $2\text{Mg}(s) + \text{O}_2(g) \rightarrow 2\text{MgO}(s) \Delta H = -1204 \text{ kJ}$
 - a. Is this reaction exothermic or endothermic? *Exothermic*
 - b. Calculate the amount of heat transferred when 2.4 g of Mg react.
-59 kJ
 - c. Will the surroundings get warmer or colder when the reaction proceeds?
Warmer
 - d. How many kilojoules of heat are absorbed when 7.50 g of MgO decomposes?
112 kJ
8. How much heat is released when 15.0 g of copper with a specific heat capacity of $0.385 \text{ J/g}^\circ\text{C}$ is cooled from 80°C to 35°C ?
289 J
9. How many kilojoules of heat are needed to raise the temperature of 10.00 kg of liquid water from 24.6°C to 46.2°C ?
982 kJ
10. If 500 J of heat are added to 100 g sample of each of the substances listed below, which will have the largest temperature increase?

<i>Solid specific heat = $0.129 \text{ J/g}^\circ\text{C}$</i>	<i>Silver specific heat = $0.237 \text{ J/g}^\circ\text{C}$</i>
<i>Copper specific heat = $0.385 \text{ J/g}^\circ\text{C}$</i>	<i>Water specific heat = $4.18 \text{ J/g}^\circ\text{C}$</i>
11. If 400.0 J of heat are added to a 150.0 g sample of water at 25.0°C , what is the final temperature of the water?
25.6^\circ\text{C}
12. What are more thermodynamically favored, *exothermic*, or endothermic reactions?
13. The heat released from the combustion of 0.0500 g of white phosphorus increases the temperature of 150.0 g of water from 25.0°C to 31.5°C . Calculate the value of the enthalpy change in kJ mol^{-1} of the combustion of phosphorus.
-2,540 kJ/mol

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